

EDITORIAL COMMENTS:

After looking at the last revision (v.6), I found that the authors have not address all the questions raised last time:

1. Define the “efficiency” in Table 1, which is not clear and actually very confusing if you look at both Fig 1 and 2 where the ethanol produced looks much smaller than the substrate consumed.

This has not addressed clearly. For example in Fig 1, about 10 gl-1 glucose consumed and about 75 gl-1 ethanol produced, so the “efficiency” would be $75/170 \times 100 = \sim 44\%$. So what is your definition of “efficiency”? If you say that “efficiency” is calculated (Fermentation efficiency of the isolates was calculated in terms of ethanol produced. page 4), this is not a definition. This has to be clarified.

2. Table 2, there seem to be a huge inconsistency of the results obtained using the same testing conditions (200 gl-1, pH 5 and 35 °) for both *P. kudriavzevii* SK1 and *S.cervisiae* MTCC 11815. There is no statistic SD provided, which seem to be not reproducible results.

200	90.62 (54)	88.83 (54)
5.0	79.4 (54)	65.4 (48)
35	65.47 (48)	53.34 (60)

I highlighted in red in your Table 2 (which are the same as I copied above). Each row above represented the same growth condition (200 gl-1, pH 5 and 35 °) with the same organism (either *P. kudriavzevii* SK1 or *S.cervisiae* MTCC 11815), but the ethanol production varied dramatically from 65.47 to 90.2 gl-1 for *P. kudriavzevii* SK1, and from 53.34 to 88.83 gl-1 for *S.cervisiae* MTCC 11815. You standard deviation values provided look very small, so how would you explain such a huge inconsistency?

This has to be clarified because the authors ignored this question in the revised version (v.6)

Based on the point #2 above, the data provided do not look reliable.

AUTHOR’S FEEDBACK:

Fermentation efficiency is calculated as follows:

$$\text{For fermentation efficiency (\%)} = \frac{\text{Actual recovery}}{\text{Theoretical recovery}} \times 100$$

Theoretical recovery (v/v) = total fermentable sugars X 0.64

Theoretical recovery (w/v) = total fermentable sugars X 0.51

For every one mole of glucose (180 g), two moles of ethanol (2X46=92 g) are produced. i.e from 1 g glucose, 0.51 g ethanol is produced.

Hence for 180 g initial sugars, theoretical ethanol(recovery)= 180X.51= 91.8 g/l

$$\text{fermentation efficiency (\%)(w/v)} = \frac{86.1}{91.8} \times 100 = 93.79\%$$

2. Standard deviation values have been added to the results in table 2. also the error in values has been corrected. We are sorry to say that previously, some error took place while copying the values and data from a different experiment was added to the table.

We hope that the manuscript meets the editor's expectations.

We are thankful to the editors and the reviewing committee for their patience and cooperation with us. Looking forward to a positive response from the editors.