

Synthesis, Properties of a New (Polymer and Chalcone)

Abstract

This study included two objectives, firstly will examine Synthesis of Polymer {poly 2- [1- phenyl -3-(4-methyl phenyl)-4, 5-dihydropyrazole-5-yl-(4-phenoxy carbonyl)] styrene} PMDPCS and structure characterized by FT-IR, ¹H-NMR spectra, Uv/Vis and XRD and secondly study was to investigate of chalcone {3-(4-(dimethylamino) phenyl)-2-phenyl-(2E) propen-1-one} (DAPPP) for absorption and fluorescence spectra under different solvents and concentrations. Also, under various concentrations, organic solvents and pump pulse energies of Nd: YAG laser, the amplified spontaneous emission (ASE) was studied. Additionally, under the same identical conditions, the amplified spontaneous emission spectra of DAPPP in solution were compared with a conventional laser dye of coumarin 503. Finally, the gain and the fluorescence quantum yield of DAPPP were determined.

Keywords

Chalcone. Solvent effect. Fluorescence. Polymer. Nd: YAG laser.

1 Introduction

Chalcones represents an important family of natural as well as synthetic organic compound. It was firstly secluded from Chinese liquorice [1]. Chalcone and its derivatives are a significant set of natural production and have been recorded to possess different biological and pharmacological effect. These chalcones had stimulated wide research interest owing to their unique structural feature, which has a carbonyl functional group in conjunction with carbon-carbon double bond. This structural arrangement gives rise to numerous chemical, physical, photophysical and biological properties [2-8]. In 2002, Yah *et al* [9] prepared several series of chalcone derivatives, which are own 90 % inhibitory activity versus *Mycobacterium tuberculosis*. First report on chalcones and related compounds as an optical material came before few decades by Kitaoka *et al* [10]. In 2003, Fabian *et al* [11] prepared 10 compounds and examined for leishmanicidal and trypanocidal activity, between that 5 compounds appeared distinguished and strong inhibitory effect on the growth of *Trypanosome cruzi* and only 2 compounds showed strong inhibitory *acripanosoma cruzi* and only two compounds showed potent inhibitory activity on the growth of *L. braziliensis* by *in vitro*. Recent reports related chalcones are most applied for different optical applications through

second harmonic generation materials in nonlinear optics and fluorescent probes for sensing of metal ions [12-21] However, the photophysical characteristics of these chalcones such as solvent environment, temperature effect and quantum yield of fluorescence were not fully studied. In the present investigation, the researcher attempted to prepared a few chalcone derivatives, which are biologically important. Also photophysical properties and amplified spontaneous emission (ASE) as well as the photochemical stability of 3-(4-(dimethylamino) phenyl)-2-phenyl-(2E)-propen-1-one (DAPPP), in different solvents and concentrations are described. ASE results were compared with conventional laser dye coumarin 503.

2 Materials and Method

2.1 Part one

2.1.1 Synthesis of Polymer PMDPCS

By free radical polymerization the polymer was prepared, the poly 2- [1- phenyl -3-(4-methyl phenyl)-4, 5-dihydropyrazole-5-yl-(4-phenoxy carbonyl)] styrene PMDPCS (0.15 gm) in presence of benzoyl peroxide C₁₄H₁₀O₄ (0.002 g) in (15 mL) tetrahydrofuran, the mix was heated with stirring at 85⁰ C for 2 hr. The polymers were collected by deposition in ice water and

washed several times and dried by diethyl ether at room temperature.

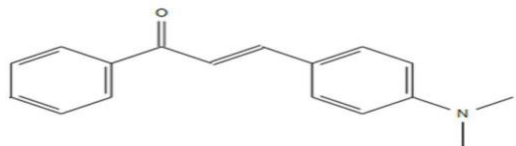
2.2 Part two

2.2.1 Synthesis of DAPPP

By condensation reaction between 4-dimethylamino benzaldehyde (0.15 mol) in the presence of NaOH (0.5 g), in (70 mL) ethanol and acetophenone (0.15 mol), the DAPPP was fabricated. The container was preserved 12 hours at room temperature and then the yield was recrystallized from ethanol and washed with distilled water and left to yield 95 %. Chalcone characterized by H^1 - NMR, C^{13} - NMR, mass and UV - vis spectroscopy.

2.2.2 Material Characterization

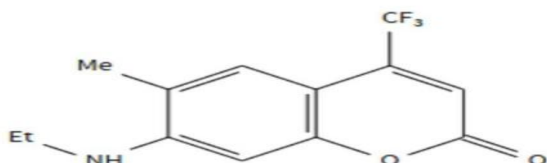
In different organic solvents, the DAPPP was resolved. Figure 1 showing the molecular structure of DAPPP. The fluorescence spectra and absorption of DAPPP in acetone were studied under a widely domain of concentrations. The spectra of these solutions were measured by a tiny quartz cuvette with an optical path length of 1 cm. A Perkin Elmer lambda 590 spectrophotometer over the range from 200 to 800 nm was used to taken the absorption spectra. A Perkin Elmer LS55 luminescence spectrofluorimeter was used to study spectra of fluorescence in the range from 200 to 900 nm, at room temperature. Fourier transform infrared spectra of PMDPCS were obtained using a Nicolet 6700 FT-IR spectrometer at room temperature in the 4000 - 400 cm^{-1} wavenumber range, with a 4 cm^{-1}



resolution.

Fig. 1 Molecular structure of DAPPP

The excitation wavelength was 400 nm. The UV laser ($\lambda = 355$ nm) was focused by a quartz plan-cylindrical lens with a focal length of 5 cm. the amplified spontaneous emission (ASE) of the DAPPP in solution was compared to that of coumarin 503 (see Fig. 2) in ethanol. The structure of the synthesized compounds was certain by IR, Mass and



elemental analysis.

Fig. 2 Molecular structure of laser dyes coumarin 503

Results and Discussion

Part One

PMDPCS was characterized by FT-IR and UV/Vis spectroscopy. Table 1 shows the FT-IR spectral data of PMDPCS. The FT-IR spectrum of PMDPCS (Fig. 3) showed the following bands at [2950, 2920] for symmetric and asymmetric stretching vibration of (-CH) aliphatic, (1070) cm^{-1} due to ν (C-O), at (1755) cm^{-1} for ν (C=O). Table 2 shows UV \Vis spectrum of prepared polymer PMDPCS, the patterns of the samples were recorded by a diffractometer (D8 Advance Bruker) using Cu-K α radiation, $\lambda = 0.15406$ nm, accelerating voltage is 40kV, scanning angle is 200 - 750 at room temperature. Figure 4 shows the XRD of PMDPCS.

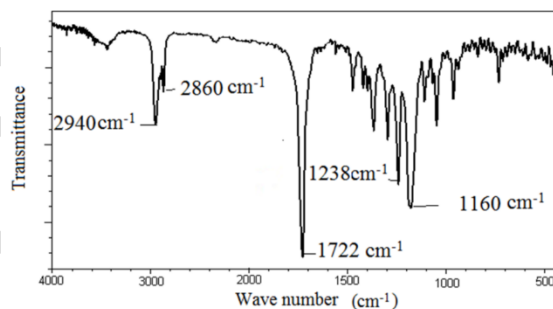


Fig. 3 FT-IR spectrum of PMDPCS

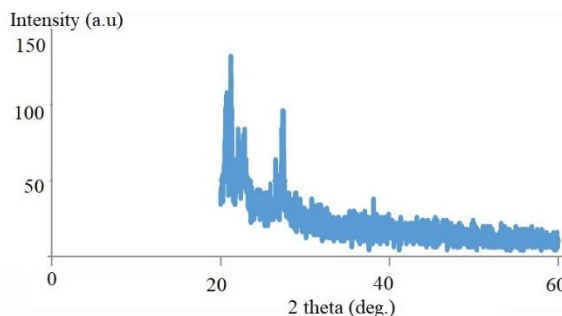


Fig. 4 XRD of PMDPCS

Part two

In different organic solvents with various dielectric constant, DAPPP was dissolved, and the concentration of these solutions was fixed at 6.5 mM. The big red shift is shown in both absorption and

emission spectra as the solvent polarity raised. Table 4 shows the influence of the solvents on the absorption and fluorescence spectra. By the above procedure, there is a significant change were observed in absorption and fluorescence spectra. As can be seen from the Fig. 4, the Stokes shift changed as a function of the dipole factor of the solvent, as defined by Lippert and Mataga et al. [23-24]. It can be seen observed that DAPPP in the solution turn strongly polar in the excited state than in the ground state. In most recent studies, the Stokes shift has a linear variation with the dipole factor, which is written by expression;

$$\nu_a - \nu_f \approx \left(\frac{(\epsilon-1)}{(2\epsilon+1)} - \frac{(n^2-1)}{(2n^2+1)} \right) \frac{(\mu_e - \mu_g)^2}{a^3 hc} \quad (1)$$

$$D_f = \frac{(\epsilon-1)}{(2\epsilon+1)} - \frac{(n^2-1)}{(2n^2+1)} \quad (2)$$

Where D_f is dipole factor, ν_a and ν_f are the absorption and fluorescence peaks in wave numbers respectively, ϵ is the dielectric constant and n is the solvents refractive index. μ_e is the dipole moment of the solute in the excited state. μ_g is representing the dipole moment of the solute the ground states, respectively, and a is the radius of the solvent cage around the solute. The results obtained from the Stokes shift as a function of dipole factor for illustrative solvents are shown in Figure 5. It can be seen that under same identical condition the DAPPP is more polar than coumarin 503. Overall, these results indicate that all these dyes exhibit large changes in the dipole moment in the excited state, but for DAPPP, the change is much greater than for coumarin 503.

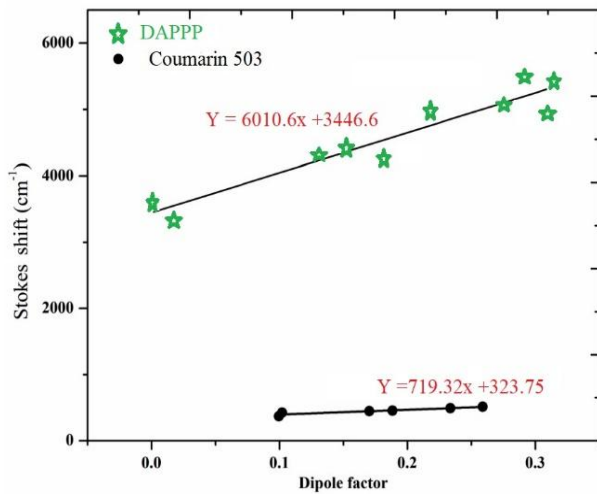
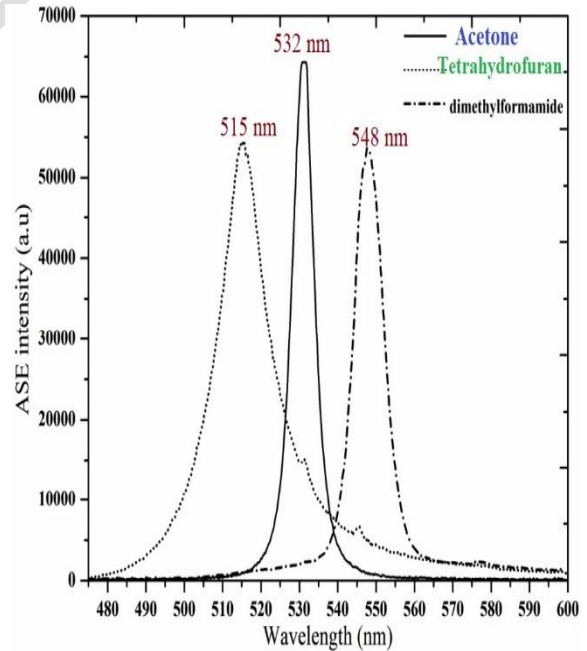


Fig. 5 Variation in the Stokes shift of DAPPP and coumarin 503.

By diluted solutions, the fluorescence quantum yields (Φ_F) of DAPPP and coumarin 503 was measured at fixed concentration for each solution 0.65 mM. The equation below illustrates how to measure the quantum yield [23];

$$\Phi_f(S) = \Phi_f(R) \frac{A_R n_S^2 \int I_S(\bar{\nu}) d\nu}{A_S n_R^2 \int I_R(\bar{\nu}) d\nu} \quad (3)$$

where the indices S and R refer to the sample and reference, respectively, and the integral over I represent the area under the fluorescence spectrum. A is the optical density, and n is the refractive index of the solvents. Tables 2 and 3 presents the results obtained from the analysis of the quantum yields of fluorescence (Φ_F) for DAPPP and coumarin 503. From the data, it can be seen that under identical conditions the Φ_F of DAPPP is higher than that of coumarin 503. Under high power excitation, present work study the ASE characteristics of DAPPP, in acetone at a concentration of 2 mM DAPPP was dissolved. Figure 6 provides the ASE spectrum (532 nm) with a narrow spectral bandwidth 6 nm (FWHM) at a pump power of 9 mJ. As shown, at 2 mM concentration the ASE peak is coinciding with the maximum of the fluorescence emission spectrum. At

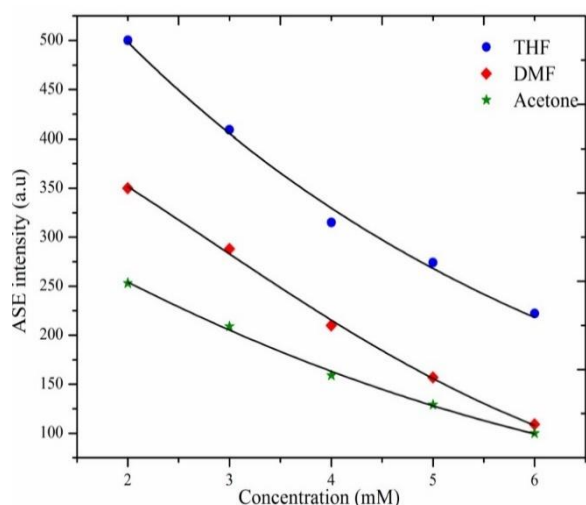


6 mM, there was no significant red shift observed.

Fig. 6. ASE spectra of DAPPP in acetone, tetrahydrofuran, and DMF.

By genetic energy meter (genetic -E), the ASE efficiency of DAPPP in acetone was measured when the input energy was 9 mJ. The energy of ASE was found to be 3 mJ. Table 2 provides that the ASE efficiency depends on the solvent. At a pump energy of 9 mJ, concentration of 2 mM and under identical conditions; the ASE spectra of DAPPP in different solvents having different polarities shown in Table 2. Fig. 6 shows the emission wavelength increasingly red shifted, when the dielectric constant increased.

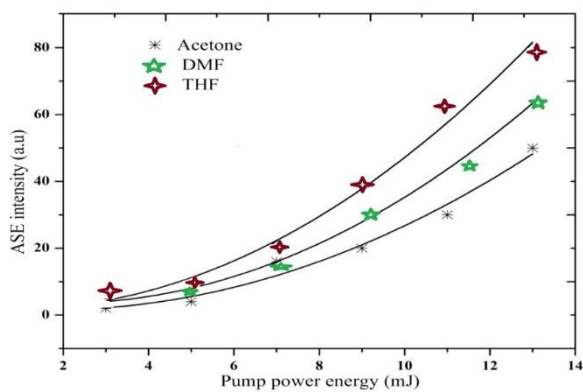
Table 3 provides the results obtained and compared with coumarin 503 as standard material. The variation in the ASE intensities of DMAPPP as a function of the concentration (2 to 6 mM) for acetone, dimethylformamide (DMF), and tetrahydrofuran (THF), and the pump power was 9



mJ. it was found that, when the concentration increased, the intensity of the ASE decreased for each solution, shown in Fig. 7.

Fig. 7 ASE intensities of DAPPP in DMF, acetone, and THF.

Under identical condition, Fig. 8 shows the comparison of the ASE intensity of DMAPPP dissolved in different solvents DMF, acetone and THF and these solvents have been chosen because they give high ASE intensities. At fixed concentration 2 mM for each solution, and the pump power was taken from 3 to 15 mJ, it was found that as the pump power was increased, the intensity of



DAPPP in acetone increased slowly, while in THF the intensity increased quickly.

Fig. 8 ASE intensities of DAPPP in DMF, acetone, and tetrahydrofuran (THF).

To compare the difference between the photochemical stabilization of the DAPPP in DMF at concentration of 2 mM to that of coumarin 503 in ethanol, UV laser of Nd: YAG was pump for this solutions with pulse energy 9 mJ and 10 Hz. The ASE intensity of DAPPP remained constant after 30K pulse, while the coumarin 503 drooped 45 % of the initial intensity after 3 K pulses, and disappeared completely at 7 K pulses as shown in Fig. 9.

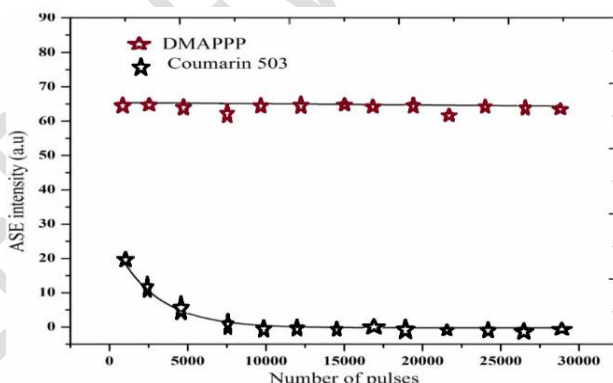
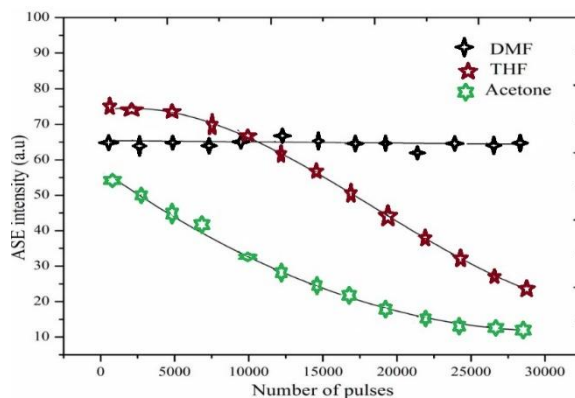


Fig. 9 The photochemical stability of DAPPP in DMF and coumarin 503 in ethanol.

By the above procedure, in different solvents and under the same operating conditions, the photochemical stability of DAPPP was studied and after 30 K pulses, the ASE intensity of THF and acetone was dropped to 35% and 22%, respectively



as seen in Fig. 10.

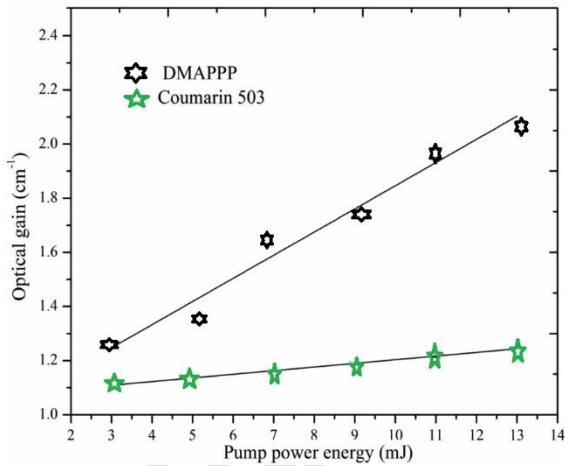
Fig. 10 The photochemical stability of DAPPP in three different solvents.

In DMF at concentration of 2 mM, the optical gain of DAPPP was measured, whereas, Coumarin 503 in ethanol was prepared. Additionally, the pump pulse energy was varied from 3 mJ to 15 mJ and the ASE intensity was measured for two lengths of excitation $L_1 = 0.5$ cm and $L_2 = 0.3$ cm [25];

$$\frac{I_1}{I_2} = \frac{e^{\gamma L_1} - 1}{e^{\gamma L_2} - 1} \quad (4)$$

where L is the length of excitation and γ is the optical gain.

The results show that the optical gain of DAPPP in DMF is higher than the coumarin 503 in ethanol at a concentration of 0.6 mM as shown in Fig. 11. The results of second objective of this study are largely consistent with the findings of A.O. Elzupir *et al.*



[26].

Fig. 11 Optical gains vs. pump pulse energy for DAPPP in DMF and coumarin 503 in ethanol

Table 1 FT-IR spectra data of compound

Compound	ν (C – H) aliphatic cm^{-1}	ν (C – H)aromatic cm^{-1}	ν (C = C) cm^{-1}	ν (C – N) cm^{-1}	ν (C = O) cm^{-1}	Other band ν (C – O) cm^{-1}
PMDPCS	2950 2920	3060	1595	1224	1755	1070

Table 2 Data of the impact of the solvents on the absorption and fluorescence spectra

Solvent	ϵ	λ_{max}			ϕ_F	η
		A	F	ASE		
Methanol	32.5	420	530	-----	0.02	----
Ethanol	24.2	419	529	547	0.35	3.8
Acetone	20.5	406	517	533	0.85	30
Toluene	2.6	405	472	-----	0.22	----
Benzene	2.23	406	475	-----	0.26	----
tetrahydrofuran	7.55	402	504	512	0.98	45

ϵ is the dielectric constant, spectral and ASE properties: λ_{max} (nm) for absorption (A), F is the fluorescence and ASE; ϕ_F ; ASE efficiency η (%) of DAPPP.

Table 3 Coumarin 503 in different solvent.

Solvent	ϵ	λ_{max}			ϕ_F	η
		A	F	ASE		
Methanol	32.7	394	490	500	0.42	9.6
Ethanol	24.5	393	492	502	0.37	16
Acetone	20.5	389	489	496	0.53	14
Toluene	2.6	378	455	465	0.21	7
Benzene	2.23	377	449	459	0.26	9
tetrahydrofuran	7.55	384	476	485	0.40	11

ϵ is the dielectric constant, spectral and ASE properties: λ_{max} (nm) for absorption (A), F is the fluorescence and ASE; ϕ_F ; ASE efficiency η (%) of coumarin 503 in different solvent.

Conclusion

In the present work, under the influence of concentrations, solvent environments, and laser power excitation levels, the spectral and ASE behaviors of DAPPP and coumarin 503 were studied. Also, Quantum yield, the Stokes shift and ASE were studied. One of the more significant findings to emerge from this study is that the DAPPP showed large the Stokes shift, and showed also a strong brightness, and green ASE. This study has succeeded to building heterocyclic compounds as monomers to be polymerized.

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